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chapter 12: coordination chemistry iv: reactions and mechanisms. 12.1. $[\text{Cr}(\text{H}_2\text{O})_6]^{2+}$ is labile, and has 4 unpaired electrons, with 1 in the anti-bonding eg orbital. Occupation of this orbital renders substitution easier by leading to relatively weak chromium(II)- aqua ligand bonds. $[\text{Cr}(\text{CN})_6]^{4-}$ is inert.

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