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this page, first find the molar mass of the compound. 1. What is the mass of 9.45 mol of aluminum oxide? mol A ILO + 3 (I A GO q 63,5ù?- 2. What is the mass of 4.52 x

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A B; percent composition: a description

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of the relative amounts of each element in a compound: empirical formula: the lowest whole- number ratio of the atoms of the elements in a compound

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25g - 6.1g = 18.9g excess of H₂

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←ANSWER *Notes: In problem B, I used 28g of N_2 again because it is the limiting reactant -- the answer I got between 28g and 25g showed that 28g had the SMALLEST number (34.04g) versus 25g's answer of 140.8g. Keep in mind that the limiting reactant is always the least.

Chapter 9: Chemical Quantities

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Chemical Quantities THE MOLE AND
QUANTIFYING MATTER 10.1 The Mole: A
... The quantities 1 mol and 22.4 L can
be used in conversion factors that
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10.2, answer the following questions.
The Mole-Mass Relationship 5.

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Chemistry Chapter 10 Chemical Quantities Test Answers

CHAPTER 10: Chemical Quantities

BASICS: • The basic unit that is used to
determine the amount of a chemical

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substance is called a mole • A mole(mol) of a substance is equivalent to 6.02×10^{23} particles of that substance • The mole was founded by a scientist named Avagadro, and he decided to use the

Chapter 10 Chemical Quantities Review Answers

Chapter 10 Chemical Quantities Study

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CHAPTER 10: Chemical Quantities

BASICS:

- The basic unit that is used to determine the amount of a chemical substance is called a mole
- A mole(mol) of a substance is equivalent to 6.02×10^{23} particles of that substance
- The mole was founded by a scientist named

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Answer Key

Avagadro, and he decided to use the

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