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1/T (K<sup>-1</sup>) 0.00180 0.00174 0.00159  
0.00150 0.00143 0.00128 ln k -10.02  
-8.92 -5.98 -4.26 -2.74 0.29 Gradient =  
-19897 = -E a 0 R E a = 19897 x 8.31 =  
165000 J mol<sup>-1</sup> = 165 kJ mol<sup>-1</sup> 6  
Temperature ( C) 20 50 100 150 200  
250 k (mol<sup>-1</sup> dm<sup>3</sup> s<sup>-1</sup>) 0.0147 0.0675  
0.497 2.29 7.61 20.1 1/T (K<sup>-1</sup>) ln k  
Gradient = -4812 = -E a R

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TASK 1 - Bronsted-Lowry acids & bases  
1 Acid = H<sub>2</sub>O, base = NH<sub>3</sub>  
2 Acid = HCl, base = H<sub>2</sub>O  
3 Acid = HCOOH, base = KOH  
Chemsheets A2 1156 (QC - Kp A)  
© wwwCHEMSHEETScouk 3-Mar-2018  
Chemsheets A2 1156 Nitrogen reacts with hydrogen as shown:  
 $3\text{H}_2(\text{g}) + \text{N}_2(\text{g}) \rightleftharpoons 2\text{NH}_3(\text{g})$   $\Delta H = -76 \text{ kJ mol}^{-1}$   
100 moles of

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Formula Sketch Shape Bond angles  
Metal oxidation state Co-ordination  
number \*CHO I 1) Draw the two

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geometric isomers of square planar complex

## **TASK 1 - Drawing complexes**

reacting mass calculations 1 1)

Aluminium is extracted from aluminium oxide as shown. Calculate the mass of aluminium that can be formed from 1020 g of aluminium oxide.

## **REACTING MASS CALCULATIONS 1 - chemsheets.co.uk | 1pdf.net**

initial moles 1 1 1 change in moles -0.8  
+0.8 +0.8 equilibrium moles 0.2 1.8 1.8  
7)  $\text{CH}_4(\text{g}) + \text{H}_2\text{O}(\text{g}) \rightleftharpoons \text{CO}(\text{g}) + 3\text{H}_2(\text{g})$   
initial moles 1 1 1 1 change in moles  
-0.2 -0.2 +0.2 +0.6

## **EQUILIBRIUM QUANTITIES 1 - WordPress.com**

1) Increase the temperature Number of particles with E<sub>a</sub> more same less  
2) Remove half the molecules Number of particles with E<sub>a</sub> more same less  
Energy  
4) Reduce the volume of the container Number of particles with E<sub>a</sub>

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3) Add a catalyst Number of Energy particles with  $E_a$  more same  $E_a$  Energy  
6) Add an inert gas more  $E_a$  less Energy

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Chemsheets A2 029 9 IR TASK 2 1) The IR spectra of six compounds are shown. The compounds are: butanoic acid  
butanone but-3-en-1-ol

## **Chemsheets A2 029 (Spectroscopy) - Weebly**

1) Propene reacts with HBr to form H. H reacts with sodium hydroxide to form I, and I reacts with warm acidified potassium dichromate (VI) to form J. The infra-red spectra of H, I and J are given below, but it does indicate which is - which.

## **IR TASK 1 - Weebly**

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Subscribe for many more exercises with answers. TASK 1 - Bronsted-Lowry acids & bases  
1 Acid =  $\text{H}_2\text{O}$ , base =  $\text{NH}_3$   
2 Acid =  $\text{HCl}$ , base =  $\text{H}_2\text{O}$   
3 Acid =  $\text{HCOOH}$ , base =  $\text{KOH}$

## **Chemsheets A2 009 (Acids & bases) ANS.pdf**

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Example 1 Imagine a reaction where R reacts with S to make some products:  $R + S \rightarrow \text{products}$  If we compare experiments 1 and 2: Initial [R] Experiment mol dm<sup>-3</sup> 1.0 2.0 Initial [S] mol dm<sup>-3</sup> 9.0 1.0 Initial rate '2.0 The concentration of R has been doubled, but the concentration of S remains the same. The rate has doubled.

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3) In an experiment, 1.56 g of propan-1-ol (CH<sub>3</sub>CH<sub>2</sub>CH<sub>2</sub>OH) was completely burned in air. The heat evolved raised the temperature of 0.250 dm<sup>3</sup> of water from 292.1 K to 339.4 K. Use this data to calculate the enthalpy of combustion of propan-1-ol (the specific heat capacity of water is 4.18 J g<sup>-1</sup> K<sup>-1</sup>).

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introduction 1) electron structure &  
definition of transition metals 1s 1s

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A rate is a measure of how some property varies with time. Speed is a familiar rate that expresses the distance traveled by an object in a given amount

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of time. Wage is a rate that represents the amount of money earned by a person working for a given amount of time.

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