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11-2 Name Symbol Molar Mass carbon  
atom C 12.01 g/mol oxygen atom O  
16.00 g/mol oxygen molecule O<sub>2</sub> 32.00  
g/mol potassium nitrate KNO<sub>3</sub> 101.11

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g/mol calcium carbonate  $\text{CaCO}_3 \cdot 2\text{H}_2\text{O}$   
136.13 g/mol dihydrate Molar mass of  
carbon dioxide ( $\text{CO}_2$ ) 12.01 g/mol + 2 ×  
16.00 g/mol = 44.01 g/mol Molar mass  
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molar mass of a gas.

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### **Class/Home Assignments**

The chemical analysis of caffeine shows



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that it contains 49.18% carbon, 5.39% hydrogen, 28.65% nitrogen, and 16.68% oxygen by mass, and its experimentally determined molar mass is 196 g/mol.  
Given: percent composition and molar mass

### **Chapter 11.2: Empirical and Molecular Formulas - Chemistry ...**

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formula mass work, Work mole problems  
name, Mole conversions work 1.

### **Molar Mass 11 3 Worksheets - Learny Kids**

Worksheet 11.1 KEY Molar Mass  
Calculations A mole is a standard unit of  
measurement for amount of a  
substance. For an element, one mole is

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equivalent to the mass of that element as it is listed on the periodic table, and thus represents its molar mass.

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Answer Key - If you find a template that you want to use, begin customizing it and you may also to open it! You will discover others call for a premium account and a number of the templates are free to use. Despite a template that is superior you may not have a handle on where to begin.



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answer key, Ws molar mass, Mole  
calculation work.

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Calculating Molar Mass (Honors  
Chemistry) Chemical Name Chemical

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Formula Atom, Molecule, or Formula Unit  
Molar Mass (g) 1. )Lead (IV) sulfite  $\text{Pb}(\text{SO}_3)_2$  327.2  
2. Dinitrogen pentoxide  $\text{N}_2\text{O}_5$  108.0  
3. Copper (II) chlorate  $\text{Cu}(\text{ClO}_3)_2$  367.3  
4. Cobalt (II) nitrite  $\text{Co}(\text{NO}_2)_2$  231.0  
5. Hydrogen gas  $\text{H}_2$  2.0  
6. Hydrochloric Acid  $\text{HCl}$  36.5  
7. Iron ...

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## **Answer Key Calculating Molar Mass**

What is the mass of  $1.25 \times 10^{23}$  sodium ions? Worksheet – Mole Conversions

Name: Show all work utilizing dimensional analysis wherever possible. Include all units and account for significant figures. A. What is the molar mass of: (make sure to show clearly defined and complete work here) 1. H<sub>2</sub> 1

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. mol H x 2.02 g H = 2.02 g H<sub>2</sub> 1 mol H<sub>2</sub>.  
2.

### **Worksheet - Mole Conversions**

Example 11. How many grams of NH<sub>3</sub> will be produced when 33.9 mol of H<sub>2</sub> are reacted according to this chemical equation?. N<sub>2</sub> (g) + 3 H<sub>2</sub> (g) → 2 NH<sub>3</sub> (g). Solution. The conversions are the

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same, but they are applied in a different order. Start by using the balanced chemical equation to convert to moles of another substance and then use its molar mass to determine the mass of the final ...

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