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An insulated container of negligible mass holds 0.600 kg of water at 45.0°C. You put a 0.0500-kg ice cube at - 15.0 o C in the water (Fig. 20.23). (a) Calculate the final temperature of the water once the ice has melted. (b) Calculate the change in entropy of the system. SOLUTION GUIDE. IDENTIFY and SET UP. 1.

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pages 541–545 page 545 1. Charged Objects After a comb is
rubbed on a wool sweater, it is able to pick up small pieces of
paper. Why does the comb ... Solutions Manual Author:

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ACQUISITION AND DEPRECIATION physics matters DISCUSSION
QUESTIONS SOLUTIONS physics matters Discuss which of the
following should be included in the cost of equipment: (a)
installation charges, (b) freight charges, (c) cost of building
foundations, (d) new parts needed to replace those damaged
while...

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$D = d \cdot 20 = (s \cdot 2) \cdot 20 = s \cdot 40 = (4.0 \text{ cm}) \cdot 40 = 25 \text{ cm}$ REASONING For convenience, we can assign due east to be the positive direction

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and due west to be the negative direction. Since all the vectors point along the same east-west line, the vectors can be added just like the usual algebraic addition of positive and negative scalars.

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