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Identify the missing

particle in the following

nuclear reaction: ${}_{99}^{43}\text{Tc} \rightarrow$

$\underline{\hspace{2cm}} + {}_{0}^{-1}\text{e} ?$

${}_{44}^{99}\text{Ru} ?$

${}_{99}^{43}\text{Tc} ?$

${}_{42}^{99}\text{Mo} ?$

${}_{99}^{43}\text{Es}$; What
is the effect of this

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decay? ${}_{99}^{43}\text{Tc} \rightarrow {}_{99}^{44}\text{Ru} + {}_{-1}^0\text{e}$? a
neutron is converted to
a proton ?

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Chapter 10 of Physical Science: Nuclear Chemistry. This is a review of the vocab, Terms in this set (17) radioactivity. the process in which an unstable atomic

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nucleus emits charged particles and energy.

radioisotope

(radioactive isotope)

any atom containing an unstable nucleus.

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Answers natural,
humans can't control it,
stop it, or slow it down.

Transmutation: when
atoms are hit or
bombarded by
particles a reaction

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may occur. Chapter 22

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Chemistry Chapter 22

nuclear chemistry.

Mass defect. Nucleons.

Nuclide. Radioactive

decay. The difference

in mass of an atom and

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ANSWER Answer the

following questions in

the space provided. 1.

Match each of the

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with the process(es) to

which they apply, using

one of the choices

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SECTION 4 SHORT

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following questions in
the space provided. 1.

Match each of the
following statements
with the process(es) to
which they apply, using
one of the choices
below: (1) fission only
(3) both fission and
fusion

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Chapter 21. Nuclear
Chemistry.

Introduction; 21.1

Nuclear Structure and
Stability; 21.2 Nuclear
Equations; 21.3

Radioactive Decay;

21.4 Transmutation
and Nuclear Energy;

21.5 Uses of

Radioisotopes; 21.6

Biological Effects of

Radiation; Appendix A:

The Periodic Table;

Appendix B: Essential

Mathematics; Appendix

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21.2 Nuclear Equations - Chemistry

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□ Review Skills 18.1

The Nucleus and

Radioactivity □ Nuclear

Stability □ Types of

Radioactive Emissions

□ Nuclear Reactions

and Nuclear Equations

□ Rates of Radioactive

Decay □ Radioactive

Decay Series □ The

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Effect of Radiation on
the Body 18.2 Uses of
Radioactive

Substances □ Medical
Uses □ Carbon-14

Dating □ Other Uses for
Radioactive Nuclides

18.3 Nuclear Energy □
Nuclear Fission and

Electric Power Plants □
Nuclear Fusion and the
Sun

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HW: questions from
review section 5.1: 1-5

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(p132); review section

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5.2 (p142): 2,3,5,7.

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Density:and

temperature

(Bozeman: Properties

of Matter: density and

gas laws) * lab: density

measurement *

lab_density

measurement (Honor)

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conversion.

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matter. Chemical vs

Physical properties and

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changes ...

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Nuclear chemistry is the study of reactions that involve changes in nuclear structure. The chapter on atoms, molecules, and ions introduced the basic idea of nuclear structure, that the nucleus of an atom is composed of protons and, with the exception

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of $1\ 1\text{H}$, $1\ 1\text{H}$,

neutrons. Recall that

the number of protons

in the nucleus is called

the atomic number (Z)

of the element, and the

sum of the ...

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20.5 Applied Nuclear Chemistry. Learning Objective. To understand how nuclear reactors function. The ever-increasing energy needs of modern societies have led scientists and engineers to develop ways of harnessing the energy released by nuclear reactions. To date, all practical applications of nuclear power have been

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fission ...
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